**Chemistry 20 – Lesson 5**

**The mole**

**/58**

**Part A**

checkmark1.



checkmark

/2

checkmark2.



checkmark

/2

checkmark3.



checkmark

/2

checkmark4.



checkmark

/2

checkmark5.



checkmark

/2

checkmark6.



checkmark

/2

checkmark7.



/2

checkmark

checkmark8.



checkmark

/2

checkmark9.



checkmark

/2

checkmark10.



checkmark

/2

checkmark11.



checkmark

/2

checkmark12.



checkmark

/2

**Part B**



**checkmark**1.



checkmark

/3

checkmark

**checkmark**

checkmark2.



/3

checkmark

**checkmark**3.



checkmark

/3

checkmark

**checkmark**4.



checkmark

/3

checkmark

**checkmark**5.



checkmark

/3

checkmark

**checkmark**6.



checkmark

/3

checkmark

**checkmark**7.



checkmark

/3

checkmark

8.



**checkmark**

checkmark/3

checkmark



**checkmark**9.



**checkmark**

checkmarkcheckmark/7

checkmarkcheckmark

checkmark

∴70 g of helium contains a greater number of moles than 250 g of water

10. If we calculate the molar mass of the gas we can use the periodic table to identify it.



checkmark/3

checkmark

checkmark

The inert gas with a molar mass of 84 g/mol is **krypton**.